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Study Guide. 1.

Terms to know: •

Electrode

Potential/Voltage:

The amount of energy per unit of charge that a half-reaction either requires or releases (typically in units of Volts) • Standard Hydrogen Electrode (SHE): A platinum electrode that

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Provides a surface
for the following
reaction: $2\text{H}^+(\text{aq}) + 2\text{e}^- \rightarrow \text{H}_2(\text{g})$ (1M)

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Multiple Choice
Identify the choice~~

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that best completes
the statement or
answers the
question. ____ 1.

Assuming the
following reaction
proceeds in the
forward direction,
$$\text{Cu}^{2+}(\text{aq}) + \text{Ca}(\text{s})$$
$$\text{Cu}(\text{s}) + \text{Ca}^{2+}(\text{aq})$$
a.

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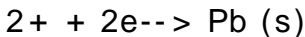
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How to find K from
 E° cell . 1. To

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potentials are
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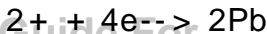
properties, meaning
that they don ' t
depend on the
amount of material.

For example: Pb.



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$E^\circ = -0.125 \text{ V}$



(s) $E^\circ = -0.125 \text{ V}$

Standard conditions
for electrochemical
cells: o Pure solids
and liquids

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~~Rachel ...~~

SAHOTA ELECTR

OCHEMISTRY

STUDY GUIDE-

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21 U 9 42. b)

Calculate the initial

voltage for the

electrochemical cell

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in part a). (1 mark)

c) Which two metals from the above list would produce an electrochemical cell with the smallest initial voltage? (1 mark) U 9 43. U 9 44. Consider the following electrochemical cell:

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electricity to flow in

a voltaic cell, the

two half-cells must

be. Choose one

answer. a.

connected to a dry

cell. b. in the same

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Answers. c. Study completely isolated from one another. d. connected by a wire and a porous barrier. d ...

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Multiple Choice -

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ALL. THE.

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QUESTIONS in this booklet. These are actual Provincial Exam questions!

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CHEMISTRY Notes

E Electrochemistry

is the study of

reactions in which

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charged particles
(ions or electrons)
cross the interface
between two
phases of matter,
typically a metallic
phase (the
electrode) and a

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Electrochemistry
Notes Calculating
Electrochemical
Potential of a Cell

- 1) Write two $\frac{1}{2}$ reactions written as reductions 2) Compare E° 's – – the reaction with the highest E° is reduced, the other is oxidized 3) Flip the oxidized

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reaction, or change the sign of E 4)

Balance/add the

reaction 5) $E_{\text{rxn}} =$

$E_{\text{reduced}} + E$

E_{oxidized}

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~~Notes~~

Electrochemical

Cell: An

electrochemical cell

is a device in which

a particular overall

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redox reaction
occurs. Specifically,
the
oxidation/reduction
half-reactions are
physically
separated into ...

~~Devise an
electrochemical cell
in which the ...
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May 2, 2020 By C.
Hall 1 Comment.

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Electrochemistry is a branch of physical chemistry (we have previously analyzed the best general books for physical chemistry), which is important enough that it deserves especial attention. It studies the movement of chemical charges, electrons, anions,

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Answers, both in solution or at the surface of electrodes. To study electrochemistry, you must have a basic understanding of physical chemistry, and math.

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Electrochemistry is the study of chemical processes that cause electrons to move. This movement of electrons is called electricity, which can be generated by movements of electrons from one element to another

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in a reaction known as an oxidation-reduction ("redox") reaction.

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